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Solution: OH OH OH 0.50 mmol mL ??

?=x = nc V ×25.0 mL OH n ? =12.5 mmol

HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = =

n V 0.10 mmol HCl(aq) /mL V =120 mL

2. (a) Given: V HCl(aq) = 25.00 mL;

[HCl(aq)] = 0.350 mol/L; V NaOH(aq) =

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5.0 mL;  $[\text{NaOH}(\text{aq})] = 0.500 \text{ mol/L}$ ;  
Required:  $[\text{H}^+(\text{aq})]$ , pH Analysis: Before  
titration, amount of  $\text{H}^+$  is:  $\text{H}(\text{aq})$   $\text{H}(\text{aq})$   
 $\text{H}(\text{aq}) 0.350 \text{ mmol mL} \quad ++ \quad += \times = \text{nc V}$

Section 8.7: Acid–Base Titration Tutorial  
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Solution:  $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.040)(0.010)}{(0.14)^2} = 0.020$

Statement: The value of  $Q$  is 0.020.  $Q$  is equal to  $K$ , so the reaction is at equilibrium. (b) Given:  $[HI(g)] = 0.20$

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mol/L;  $[H_2(g)] = 0.15 \text{ mol/L}$ ;  $[I_2(g)] = 0.090 \text{ mol/L}$ ;  $K = 0.020$  Required:  $Q$ ;  
direction of reaction to reach equilibrium  
Solution:  $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.15)(0.090)}{(0.20)^2} Q = 0.34$

Section 7.5: Quantitative Changes in  
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