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The reaction for  
the decomposition  
of hydrogen

peroxide is  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ .

List three  
ways to speed up  
the rate of  
decomposition. For

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each one, briefly explain why it is effective, based on collision theory.

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